

Binding ability of lacosamide to phosphatidylcholine lipids with cholesterol effects

Lakozamid'in fosfatidilkolin lipitlerine bağlanma kapasitesi ve kolesterolün etkisi

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ABSTRACT

Aim: In the present study, in order to obtain information about binding profile of lacosamide (LCM) with membrane lipids, which is the most important parametre that determines its availability in the central nervous system, we aimed to analyze molecular interaction of LCM with model membranes formed by well-known representative lipids of biological membranes .

Methods: The effects of varied concentrations (1-10-20 mol %) of LCM on biophysical parameters of dipalmitoylphosphatidylcholine multilammelar vesicles in the absence and presence of cholesterol (10 mol %) were studied by differential scanning calorimetry and fourier transform infrared spectroscopy.

Results: According to the data obtained in all concentrations LCM decreased main transition temperature and enthalpy with broaden transition curve. However, LCM at 1 mole % affected such values mostly in comparison with 10-20 mol% concentrations. With the addition of CHO to 20 mol % of LCM sample, these parameters were significantly lowered. Moreover, LCM at high concentrations prominently increased both order of lipid phases, and hydrogen binding capacity of glycerol and phosphate groups of lipids but reduced fluidity. Also, CHO was found to tend to reverse these effects of LCM except for order and fluidity parameters.

Conclusion: According to the findings obtained, in accordance with its chemical structure, LCM leads to alteration of thermotropic parameters, structures and functions of phosphatidylcholine lipids abundant in brain cell membrane, which also show the penetration ability of LCM into biological membranes. But, the degree of such behavior can be dependent on applied concentration of LCM and presence of CHO. In summary, all data may provide information for its use in therapeutic strategies of epilepsy and development of new antiepileptics.

Keywords: Lacosamide, dipalmitoylphosphatidylcholine, cholesterol, infrared spectroscopy, differential scanning calorimetry

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Öz

Amaç: Bu çalışmada amaç, Lakozamidin (LCM) merkezi sinir sistemindeki etkinliğini belirleyen en önemli parametre olan zar lipitleri ile bağlanma profili hakkında bilgi elde etmek için, biyolojik zarlarla en iyi bilinen temsili lipidleri tarafından oluşturulan model zarlarla LCM'nin moleküler etkileşimi incelemektir.

Yöntem: Bu amaç doğrultusunda, farklı konsantrasyonlarda (1-10-20 mol%) LCM'nin dipalmitoilfosfatidilkolin çok tabaklı veziküllerin biyofiziksel parametreleri üzerindeki etkileri kolesterolün (CHO) yokluğunda ve varlığında (% 10 mol) diferansiyel tarama kalorimetrisi ve fourier transform infrared spektroskopisi ile araştırıldı.

Bulgular: Elde edilen verilere göre, tüm konsantrasyonlarda LCM'nin ana geçiş sıcaklığını, ana geçiş eğrisinde genişleme ile birlikte entalpiyi azalttığı, bununla birlikte, %1 mol oranında LCM'nin, %10-20 mol konsantrasyonlarına kıyasla bu değerleri en fazla etkilediği belirlendi. Öte yandan, CHO'ün %20 mol'e eklenmesi, bu parametrelerde ölçüde azalmaya neden olduğu bulundu. Ayrıca, yüksek konsantrasyondaki LCM, gliserol ve fosfat gruplarının hidrojen bağı yapma kapasitelerini ve lipid düzenini arttırmış fakat akışkanlığı azaltmıştır. Ayrıca, CHO'nun lipit düzeni ve akışkanlık dışında LCM'nin bu etkilerini tersine çevirmeye eğiliminde olduğu bulundu.

Sonuç: Elde edilen bulgulara göre kimyasal yapısına uygun olarak LCM, beyin membranında bol miktarda bulunan fosfatidilkolin lipitlerinin termotrofik parametrelerinin, yapılarının ve işlevlerinin değişmesine yol açması LCM'nin biyolojik membranlara nüfuz etme yeteneğini göstermektedir. Ancak, bu tür davranışların derecesi, uygulanan LCM konsantrasyonuna ve CHO'nun varlığına bağlı olabilir. Özette, tüm veriler epilepsinin terapötik stratejilerinde kullanılımı ve yeni antiepileptiklerin geliştirilmesi için bilgi sağlayabilir.

Anahtar kelimeler: Lakozamid, dipalmitoilfosfatidilkolin,コレsterol, kızılıtesi spektroskopisi, diferansiyel tarama kalorimetresi

INTRODUCTION

Lacosamide (LCM) ([R]-2-acetamidoN-benzyl-3-methoxypropionamide), (UCB, Inc. Smyrna, GA), (Figure 1) is an antiepileptic drug with a molecular weight of 250.30 Da used as add-on therapy for partial-onset seizures with antinociceptive and neuroprotective activities^{1,2}. Also named vimpat, harkoseride and SPM 927, it is an analog of the amino acid L-serine. The primary mechanism by which LCM exerts its anticonvulsive effect is related to its role in the enhancement of slow inactivation of component of voltage-gated sodium channels (VGSCs) without affecting the fast inactivation, thereby stabilizing hyperexcitable neuronal membranes¹⁻⁶. LCM has amphiphilic properties, allowing the molecule to be sufficiently water soluble (0.465 mg/ml), but also lipophilic enough to cross the blood-brain barrier ($\log P = 0.02$)⁷. It has hydrogen-bond donor and acceptor count as 2 and 3, respectively. Its topological polar surface area is 67.4 A²⁸. Based on such characteristics, LCM shows favorable pharmacokinetic profile in humans, with rapid and complete absorption, low plasma protein binding (<15%), and low potential for clinically relevant drug-drug interactions^{9,10}. In addition, the drug is rapidly absorbed after oral administration, reaching maximum plasma concentrations after 0.5-4 h and it is almost completely eliminated in the urine within about 13 h¹¹.

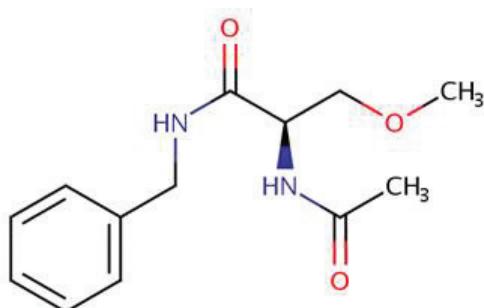


Figure 1. Chemical structure of lacosamide.

Since LCM exerts its action on VGSCs (4-5) that are integral proteins in central nervous system (CNS), it inevitably interacts with membrane lipids in order to reach target sites in CNS, and thereby binding at correct position to VGSCs like other antiepilep-

tic agents^{12,13}. For that reason, the penetration of LCM into lipids directly affect its pharmacokinetic and pharmacological properties including its CNS availability¹²⁻¹⁵ as well as its toxicity and cell resistancy as shown for different drugs^{14,15}. Thus, the evaluation of LCM in terms of the binding ability with membrane lipids may provide additional information for its partition profile, and accordingly its effective use in epilepsy treatment. However, although LCM is expected to interact with membrane lipids due to its high distribution pattern throughout the body and very low protein binding capacity^{9-11,16}, there is no published data on LCM-lipid interaction to the best of our knowledge.

Displaying a very high complexity with varied lipid, proteins and carbohydrates, biological membranes are accompanied by some disadvantages to examine drug-lipid interactions because of difficulties in the interpretation of the results. Thus, usage of model membrane systems which offer specific information is a general approach for such studies^{18,22-24}. Specifically, the composition of outer leaflet of plasma membranes of mammalian cells consists of phosphatidylcholines (PC) and sphingomyelin, which together reach to abundant amounts in most membranes, and the other main components such as cholesterol (CHO) found up to 50 mol %^{23,25}. Model systems made up of CHO and major lipid components such as PC lipids like dipalmitoylphosphatidylcholine (DPPC) have been widely used to mimic biological membranes^{13-15,21,22,24,25}.

It is very well-known that when drug molecules interact with membrane lipids, biophysical parameters of lipids such as phase transition enthalpy (ΔH), phase transition temperature (T_m), lipid order, lipid fluidity and hydrogen state of glycerol and phosphate groups are changed. By monitoring such changes in model membrane systems the information about the degree of interaction and behaviour of the drug in real membrane systems can be obtained^{14,17,26-27}. Similarly, in the present study we aimed to provide information for binding ability of LCM with PC membrane lipids, and designed our experiments on a well-established

molecular models of biomembranes^{13,18,22,24,26}, multilamellar vesicles (MLVs) formed by DPPC and CHO lipids, as similarly performed for phenytoin,¹² and gabapentin and levetiracetam¹³. Indeed, in the absence and the presence of CHO (10 mol%), the effects of LCM at different concentrations (1-10-20 mol%) on biophysical parameters of DPPC MLVs were investigated. To achieve this, Fourier transform infrared (FT-IR) spectroscopy and differential scanning calorimetry (DSC) were utilized as in earlier studies^{12,13,17-19,22,24-27}.

MATERIAL and METHODS

Chemicals

Pure active substance of LCM, DPPC, CHO and phosphate buffered saline (PBS) tablets at the highest available grade of purity were purchased from Sigma (Sigma Chemical Co, St Louis, Mo, USA).

Preparation of MLVs

MLVs containing pure DPPC, 1-10-20 mol% LCM+DPPC and 20 mol % LCM+10 mol% CHO+DPPC were prepared in accordance with the previous reports^{13,22,24,28}. Briefly, 5 mg of DPPC lipid was initially dissolved in chloroform in a round-bottom glass tube, and 5 mg DPPC was accepted as 100 mol %, which was 6.8×10^{-6} mol. Then the solvent was removed under nitrogen flow and dessicated vacuum to obtain thin dried lipid film. The films were re-suspended in PBS buffer, which was already obtained by dissolving PBS tablets in deionized water, resulting in a solution of 0.01 M phosphate buffer with 0.0027 M potassium chloride and 0.137 M sodium chloride with a pH of 7.4. MLVs were formed by vigorous vortexing the mixture for 30 min at about 60°C that is above the main phase transition temperature (Tm) of DPPC lipids. To prepare 1 mol % of LCM containing MLVs, the appropriate amount of LCM (6.8×10^{-6} mol \times 1/100 = 6.8×10^{-8} mol = 1.70×10^{-5} g) was taken from a stock solution, in which LCM was dissolved in ethanol and placed in a round-bottomed a glass tube. Excess ethanol was evaporated by nitrogen stream and then DPPC in chloroform was added. The solvents were again eva-

porated by a stream of nitrogen, and MLVs were prepared as described above. In order to produce other MLVs, the required amounts of DPPC, LCM and CHO were calculated by using the same proportioning and similar procedures.

FT-IR Studies and Analysis

Infrared spectrum of each MLVs was acquired by a Perkin Elmer Spectrum 100 FT-IR spectrometer (Perkin Elmer, Inc., Norwalk, CT, USA) equipped with a deuterated triglycine sulfate detector. The sample was placed between calcium fluoride windows with 12 μm spacer to give sample thickness. The interferograms were averaged for 200 scans at 2 cm⁻¹ resolution. Each sample were scanned between 25-60°C temperature range with increasing rate of 3°C, and at each temperature degree the samples were incubated for 5 min. All experiments were performed three times. The sample compartment in the FT-IR spectrometer was continuously purged with dry air to prevent water vapor. The spectrum of air was automatically subtracted by Perkin Elmer Spectrum One software (Perkin Elmer, Inc., Norwalk, CT, USA).

Since all samples were placed in buffer the OH stretching modes (3400-3200 cm⁻¹ and 1800-1500 cm⁻¹) from water molecules overlap the bands of interest. To eliminate the effect of water on spectra, buffer spectrum at each corresponding temperature was subtracted by using the same software.

Perkin Elmer Spectrum One software was again utilized for spectral analysis. Detailed data analysis was performed on water subtracted sample spectra. The band positions were measured from the center of weight ($0.75 \times$ peak height positions) and bandwidth value was calculated as the width at $0.75 \times$ height of the signal in terms of cm⁻¹.

DSC Studies and Analysis

DSC studies and analysis were done in accordance with the literature^{13,22,24,28}. For the sample preparation of DSC studies, MLVs containing; pure DPPC,

1-10-20 mol% of LCM+DPPC, and 20 mol%LCM+10 mol % CHO+DPPC were prepared by the same procedure used in FT-IR studies. Each MLV suspension was first encapsulated in hermetically sealed standard aluminum DSC pans by lid crimper. Then, the sample was analyzed by a Shimadzu DSC-60 Calorimeter (Shimadzu Corporation, Tokyo, Japan). During analysis, each sample was equilibrated at 25°C and then heated to 60°C at a rate of 0.5°C/min. Samples were scanned three times to ensure the reproducibility of the endotherms. Data were analyzed using TA 60 software provided by Shimadzu (Shimadzu Corporation, Tokyo, Japan). Pure DPPC samples resulted in two endothermic thermal events as shown in Figure 2. The temperature at the first peak maximum, named pre-transition, was obtained as 32°C. The second peak called main phase transition was measured at 41.11°C. The area under the main transition peak was obtained for the enthalpy (ΔH) values as performed by Türker-Kaya et al. in their study (2017)^{13,22}.

Statistical Analysis

To evaluate the differences among the samples in comparison with pure DPPC MLVs statistical significance was assessed by post-hoc Tukey HSD following One Way Analysis of Variance (ANOVA) using SPSS Statistics 19.0 program (IBM SPSS Statistics for Windows, Version 19.0. Armonk, NY: IBM Corp). Differences were considered statistically significant when $p<0.05$.

RESULTS

Figure 2 shows thermograms of DPPC MLVs with and without LCM and CHO. As demonstrated in the figure,

in pure DPPC thermogram pre-transition and main transition peaks were obtained as 32°C and 41.11°C, respectively. These values are in accordance with the literature^{13,23}. Particularly, the peak maximum of main transition, T_m value, is representative of transition of lipids from gel to liquid phase. In other words, above 41°C, DPPC lipids are in liquid crystalline phases while below that degree they are in gel phases^{13,23,26}.

As shown in Figure 2, both LCMs at all concentrations with CHO caused disappearance of pre-transition peak. Table 1 illustrates numerical comparisons of T_m and ΔH values of all samples. Table 1 shows a slight decrement in T_m but significant ($p<0.05$) decrease in ΔH values compared to pure DPPC. Indeed, according to Tukey test, there was no statistically difference among T_m of LCM and CHO groups however, ΔH values of such groups were found to be different at a significance level of 0.05. Among all applied concentrations 1 mol % of LCM was the most effective one, which caused the highest reduction in T_m (from 41.11 ± 0.19 to 39.60 ± 1.12) and ΔH (from 35.67 ± 2.70 to 24.57 ± 1.08) values, respectively. But with the addition of 10 and 20 mol % of LCM these parameters tend to become closer to pure DPPC values. Subse-

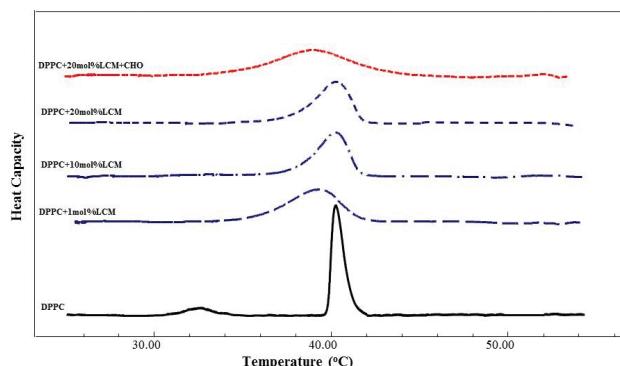


Figure 2. DSC thermograms of DPPC MLVs in the absence and presence of LCM and CHO.

Table 1. T_m and ΔH values of DPPC MLVs in the absence and presence of LCM and CHO.

Sample	T_m (°C)	P	ΔH (j/g)	P	Bandwidth (cm ⁻¹)	P
DPPC	41.11 ± 0.19		35.67 ± 2.70		22.40 ± 0.85	
DPPC+1mol%LCM	39.60 ± 1.12	0.49	$24.57^*\pm1.08$	0.001	22.19 ± 1.23	1
DPPC+10mol% LCM	40.15 ± 1.87	0.82	$25.58^*\pm2.36$	0.002	21.94 ± 1.57	0.98
DPPC+20 mol%LCM	40.42 ± 0.95	0.93	$26.05^*\pm1.67$	0.003	20.58 ± 0.92	0.30
DPPC+20mol%LCM+10mol %CHO	39.12 ± 0.65	0.25	$22.08^*\pm2.94$	0.000	20.47 ± 0.41	0.25

quently, T_m and enthalpy values of DPPC MLVs containing 20 mol % of LCM were the closest ones to pure DPPC. This may suggest that LCM effects thermotropic parameters of DPPC lipids independent of its concentration. The presence of 10 mol % CHO was found to behave quite differently, by reducing T_m and also significantly decreasing enthalpy ($p<0.05$) values mostly in comparison with the other samples.

For FT-IR spectral analysis the wavenumber values of the CH₂ asymmetric ($\sim 2925\text{ cm}^{-1}$), the C=O ($\sim 1730\text{ cm}^{-1}$) and the PO₂-2 asymmetric ($\sim 1225\text{ cm}^{-1}$) stretching were evaluated, and the bandwidth values of the CH₂ asymmetric stretching were calculated. Figure 3 illustrates FT-IR spectra of pure and 20 mol % LCM and CHO containing DPPC MLVs in the wavenumber range of 3000-2800 cm^{-1} at the liquid crystalline phase. The changes in wavenumber values of the CH₂ asymmetric, the C=O and the PO₂-2 asymmetric bands for pure DPPC, DPPC+20 mol % LCM and DPPC+20 mol % LCM+10 mol % CHO are depicted in Figure 4A, 4B and 4C, respectively. The bandwidth values of the CH₂ asymmetric of the samples are also included in Table 1. As demonstrated in Fig. 3 and Fig. 4A, significant ($p<0.05$) shifting to lower values of the CH₂ asymmetric mode was obtained with addition of LCM at 20 mol %, and CHO further lowered such impact in both gel and liquid crystalline phases. In comparison with pure DPPC, the bandwidth value of the same mode slightly decreased with the presence of LCM, and CHO enhanced this reduction (Table 1). According to Figures 4B and 4C, LCM at the highest concentration significantly ($p<0.05$) lowered the frequency values of C=O and PO₂-2 asymmetric stretching modes of DPPC lipids. The presence of CHO still caused a decrement ($p<0.05$) in such parameter, however in comparison with LCM sample CHO showed tendency to increase these wavenumber values. Post-hoc Tukey test revealed that in terms of the CH₂ asymmetric, the C=O and the PO₂-asymmetric wavenumber values LCM and CHO groups were significantly different from each other but bandwidth values of these groups did not show any difference.

DISCUSSION

Considering its low protein binding and level of distribution in body, LCM is expected to incorporate into membrane lipids. In order to provide information for this, in the absence and presence of CHO, the effects of LCM at different concentrations (1-10-20 mol%) on thermotropic properties (T_m and ΔH), hydration state of head groups, glycerol backbones, lipid dynamics (fluidity) and lipid acyl chain flexibility (ordering) of DPPC MLVs were monitored. Owing to their advantages to study such biophysical features of lipids, DSC and FT-IR spectroscopy were utilized. DSC offers measurement of thermodynamic parameters (i.e. ΔH and T_m); FT-IR spectroscopy offers data about lipid dynamics (fluidity), lipid acyl chain flexibility (ordering), hydration state. Thus, these techniques have been extensively utilized in these kinds of works^{13,17-19,22,24-28}.

DSC thermogram of pure DPPC displayed pre-transition peak around 32°C, which is related with the mobility of choline and polar head groups of lipids¹⁷ (Figure 2). The disappearance of pre-transition peak with addition of CHO and LCMs may suggest extreme broadening and/or disturbance of ripple phase. Since pre-transition peak depends on choline and head groups of lipids, hydration state of such groups directly affects periodic ripples on membrane surface^{13,24,28}. Thus, this may indicate that both LCM and CHO accumulate at the surface of bilayer by showing tendency to interact with hydrophilic parts of DPPC lipids. This behavior of CHO is in accordance with the studies on CHO-DPPC interactions^{24,25}. In addition, with regard to LCM data similar result was published related to different antiepileptics^{12,13}. Nevertheless, it is worth noting that pre-transition is extremely susceptible to any foreign molecules, perturbation of pre-transition peak may not be associated with alterations in hydration state of lipids^{12,21}. In order to bring more clarification to this issue, wavenumber values of C=O and PO₂-symmetric stretching bands were evaluated within the scope of the present study. The C=O mode is attributed to the carbonyl groups of the glycerol backbone near hydrophilic region while PO₂-band

is assigned to head groups of lipids^{13,17,22,24-27}. Due to their chemical features both LCM and CHO are expected to interact with such groups over hydrogen bonding. In detail, LCM has 2 hydrogen bond donors and 3 acceptor counts⁸ while CHO has 1 hydrogen donating and 1 hydrogen accepting group²⁹. Thus, C=O and PO₂⁻ would be moiety for LCM and CHO to form hydrogen bonds. As demonstrated in Figures 4B and 4C, significant ($p<0.05$) decrease in wavenumber values of both peaks in the presence of 20 mol % of LCM and 10 mol % of CHO, may suggest formation of hydrogen bonds between such groups^{17,22,28}. For LCM, the hydrogen bonding can occur between the C=O and the PO₂⁻ groups of DPPC lipids and either with N-H and O groups of LCM or water molecules rather than lipids groups of each other. To the best of our knowledge, even though the solubility of LCM is known⁹, such ability of LCM on lipids has not been previously reported. Same result was also obtained in our previous study for gabapentin and levetiracetam, both of which have amphiphilic character like LCM¹³. Even though in the sample containing LCM and CHO there was still decrement in wavenumber values of C=O and the PO₂⁻ modes, the degree of such reduction seemed to be less pronounced than that of LCM itself. This may have resulted from competition between LCM and CHO. Particularly, polar OH groups of CHO would make hydrogen bonding with PCs oxygen atoms in the C=O and the PO₂⁻ groups, and water^{22,24,25}. In addition, it is also probable that there might be an association between LCM and CHO. But, probability of such influence of CHO is very low because number of hydrogen bond donor and acceptor groups are 1 and 1, respectively²⁹.

For the evaluation of DSC and FT-IR numerical data we applied post-hoc Tukey test following ANOVA and we were able to compare not only pure DPPC with other samples but also LCM and CHO groups with each other. At the level of significance ($p<0.05$) there was no difference among each group in Tm and bandwidth values. However, it should be mentioned that significance level of DPPC+20 mol% LCM+10 mol% CHO sample for Tm (0.25) and for bandwidth (0.25) are higher than that of the other samples as

shown in Table 1. Even though such finding was expected to be interpreted as no difference among all samples, for model membrane studies such variations may still mean that there is interaction between LCM, CHO and DPPC lipids. This may also reflect as significant changes in ΔH values as well as the data of FT-IR study.

Figure 2 shows thermograms of all samples, broadening in the main transition curve as well as a reduction in Tm and ΔH values (Table 1) with the addition of LCM and CHO. Since main transition peak is correlated with the mobility of fatty acyl chains^{17,19}, such parameters reflect phase behavior of lipids dominated by lipid-lipid and binding of any molecules, which may be further referred to partition of LCM with CHO into DPPC lipids^{8,12,14,17}. As being marker of transition from gel phase to liquid crystalline phases for DPPC lipids, broadened main transition peak may reveal lower number of molecules undergoing transition simultaneously^{28,30}. Furthermore, decrement in Tm and ΔH may also suggest incorporation of both LCM and CHO, which may mean perturbation of normal phase transition of DPPC lipids³⁰. In other words, the total energy involved in the transition is decreased because the number of lipid molecules involved in the temperature-induced transition is reduced^{24,27,28,30-32}. This might have resulted from localization of LCM and CHO in outer hydrophobic cooperative zone of the bilayer, i.e. the region of C1-C8 atoms of the acyl chains by disturbing van der Waals interactions between acyl chains. Moreover, the decrease in Tm may also have resulted from adsorption of water molecules by hydrophilic parts of DPPC lipids. This further reduces the strength of the interactions of adjacent lipid molecules in the bilayer primarily by disruption of polar head group-head group interactions^{13,17,32} with the presence of LCM as also observed from the frequency values of C=O and PO₂⁻ bands. However, it should be stated that the effect of LCM on thermotropic behavior does not depend on its increasing concentration. In detail, the presence of 1 mol% of LCM decreases such parameters at a maximum level. This might explain that LCM at low concentration is very effective on thermotropic parameters of DPPC

lipids. With the addition of CHO, these parameters were lowered at the highest level. This might be caused by higher penetration capacity of CHO than that of LCM into DPPC bilayers. In particular, it is well-known that in addition to its four ring hydrogen bonding abilities of OH groups, it could facilitate positioning of CHO in depth of DPPC membranes. For this reason, even though 20 mol % of LCM tends to behave closer to pure DPPC thermotropic values; presence of CHO could reverse such action of LCM at high concentrations. This behavior of CHO can be supported by earlier studies in which only CHO was added into DPPC systems and, it showed similar effect such as very broad main transition, and reduction in T_m and ΔH values^{24,33,34}.

As obtained from C=O and PO₂⁻ modes, hydrogen bonding of LCM and CHO with both carbonyl and head groups of DPPC lipids may point out that both CHO and LCM have potency to alter lipid acyl chain flexibility and lipid dynamics as reported by Altunayyar et al.¹⁹. To elaborate on this, the effects of these molecules on lipid order and fluidity of DPPC MLVs were also analyzed. The frequency alterations in the CH₂ asymmetric stretching band provide information about the order/disorder state (acyl chain flexibility) of membrane lipids^{13,17,22,24,28,30,32}. As illustrated in Figure 3 and Figure 4A, LCM at 20 mol % led to significant ($p<0.05$) reduction in wavenumber values in both gel and liquid crystalline phase, which may suggest increased trans conformers within the system indicating a higher order of acyl chain packing and the rigidity of the membrane^{13,17,32}. In general, when orderliness in lipid bilayer system is observed, higher T_m and enthalpy values are expected³¹. However, the detected decrease in such parameter might be caused by strong hydrogen bonding of LCM with carbonyl and phosphate head groups of lipids. The ordering effect of LCM at high concentration may stand for its stabilizing effect on neuronal membrane to recover and/or to prevent disordered epileptic membrane systems as similarly obtained for levetiracetam¹⁷. Moreover, with the presence of CHO more decrement observed in T_m value in comparison to LCM sample may suggest that CHO further stabilizes the membrane

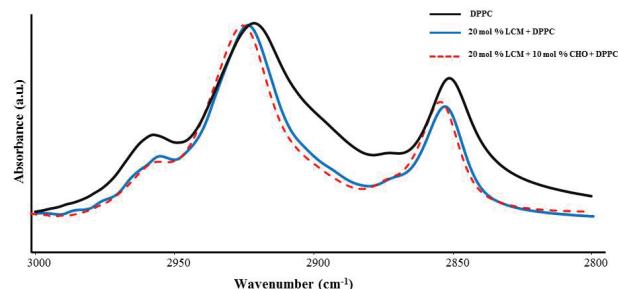


Figure 3. Representative FT-IR spectra of DPPC MLVs in the absence and presence of 20 mol % LCM and 10 mol % CHO in the wavenumber range of 3000–2800 cm⁻¹ at the liquid crystalline phase. The spectra were normalized with respect to the CH₂ asymmetric stretching mode.

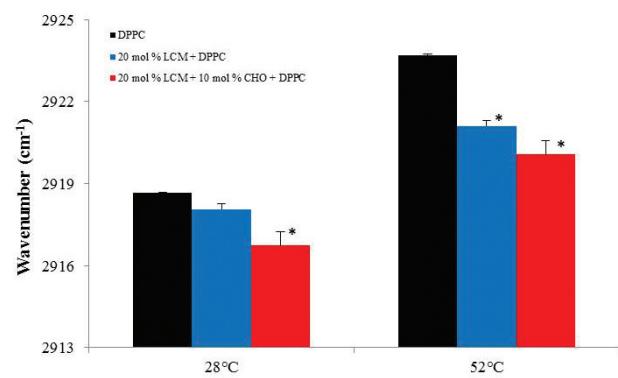


Figure 4A. The wavenumber values of the CH₂ asymmetric stretching mode for DPPC MLVs in the absence and presence of 20 mol % LCM and 10 mol % CHO in gel and liquid crystalline phases.

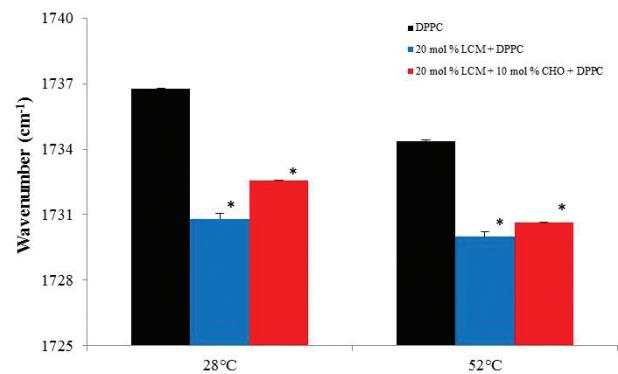


Figure 4B. The wavenumber values of the C=O stretching mode for DPPC MLVs in the absence and presence of 20 mol % LCM and 10 mol % CHO in gel and liquid crystalline phases.

system, which is also in line with the literature²⁴⁻²⁷.

Knowledge about the dynamics of DPPC MLVs is obtained from the bandwidth of the CH₂ asymmetric stretching vibration^{22,28,32}. A slight decrease in the

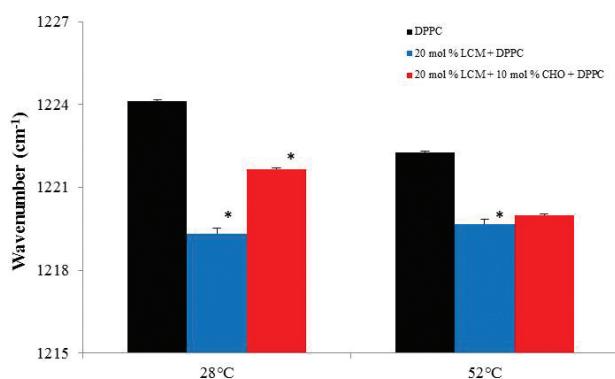


Figure 4C. The wavenumber values of the PO_2^- - asymmetric stretching mode for DPPC MLVs in the absence and presence of 20 mol % LCM and 10 mol % CHO in gel and liquid crystalline phases.

bandwidth of the CH_2 antisymmetric stretching band of DPPC MLVs in the presence of 20 mol% concentration of LCM suggests the mobility of the hydrocarbon chains showing less fluid system^{28,31,32}. The behavior of LCM on fluidity of DPPC MLVs may be related with its preventive effects on generation of action potential. This may be also associated with its use against neuropathic pain³⁵. When CHO was added to the lipid system, lowered bandwidth value of CH_2 asymmetric stretching mode may indicate that CHO makes the system less fluid. In principle, when applied to pure DPPC MLVs, CHO decreases fluidity in gel phase but increases it in liquid crystalline phases^{24,32}. However, in our case, the observed decrement in fluidity of DPPC lipids with the addition of CHO might be resulted from additive reducing impact of LCM on fluidity.

As being the main limitation of the current study, the experiments were conducted on model membrane systems containing two membrane lipids which may mimic biological membrane systems. In order to bring about more clarification on the partition profile of LCM in biological membranes, further studies should be performed on isolated lipids from animal tissues. Furthermore, in this study, the model membrane system MLVs, which is accepted as one of the representative models of biomembranes, and the utilized techniques, FT-IR spectroscopy and DSC, have been widely used to investigate drug-lipid interactions. Even though these techniques and model membrane systems provided general information for

binding ability of LCM with membrane lipids, in order to obtain more solid data this study can be extended by using additional model membrane systems such as large unilamellar vesicles and bilayers formed by different lipids. Additionally, other spectroscopic methods giving more specific information like electron spin resonance and nuclear magnetic spectroscopy can be performed.

CONCLUSION

Our results showed that LCM at applied concentrations (1-10-20 mol%) can alter thermotropic parameters of DPPC lipids without depending on increasing concentrations. Investigated concentrations of LCM incorporate into DPPC MLVs and localize within acyl chains of lipids. Additionally, LCM forms strong hydrogen bonding with hydrophilic parts of lipids. Moreover, it also decreases fluidity but increases order. Summarized findings show that binding capability of LCM is highly correlated with its chemical structure. In addition, CHO further reduced thermotropic parameters, fluidity and hydration state of hydrophilic groups but increased order of DPPC lipids. These results provide ground for better understanding of behavior of LCM within biological membrane systems, especially those containing primarily PCs and CHO which are widely found in brain cell membrane.

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